

Synthesis and characterisation of the aminocarbene complexes $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{NHR}^2)(\text{CH}_2\text{R}^1)\}]^+\text{BF}_4^-$ [$\text{R}^1 = \text{H, Me}$ or Pr ; $\text{R}^2 = \text{H, Me, Et, CHMe}_2, \text{CH}_2\text{Ph, CH}(\text{Me})\text{Ph, CH}_2\text{CH}=\text{CH}_2$ or $\text{CH}_2\text{CH}_2\text{OH}$]

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A number of secondary aminocarbenes of the type $(RS)-[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{CH}_2\text{R}^1)(\text{NHR}^2)\}]^+\text{BF}_4^-$ were synthesised in very good yield [66–99%, $\text{R}^1 = \text{H, Me}$ or Pr ; $\text{R}^2 = \text{H, Me, Et, CHMe}_2, \text{CH}_2\text{Ph, CH}(\text{Me})\text{Ph, CH}_2\text{CH}=\text{CH}_2$ or $\text{CH}_2\text{CH}_2\text{OH}$]. The absolute configuration of one of the carbenes was established by a single-crystal X-ray diffraction experiment. The chemistry of the carbene complexes was investigated.

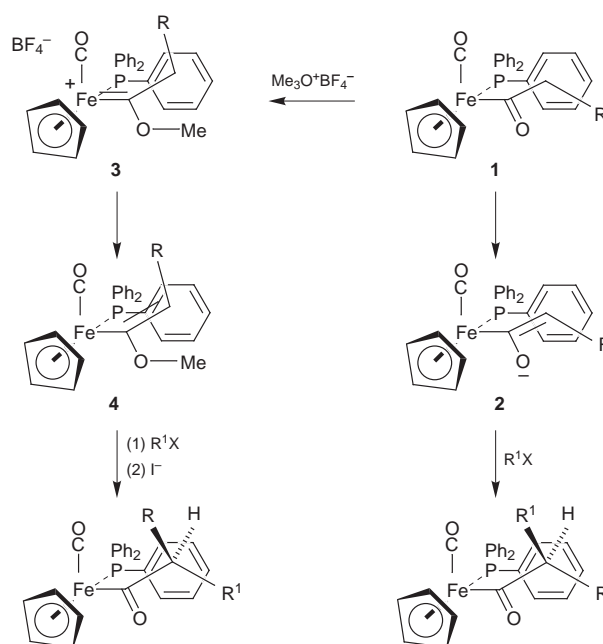
Previously we have established the fragment $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)]^{1-3}$ as a versatile chiral auxiliary for asymmetric organic synthesis. This iron chiral auxiliary induces high stereoselectivity in the reactions of a wide variety of attached ligands. For example, deprotonation of the acyl ligand in $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\text{COCH}_2\text{R}]$ **1** generates the *E*-enolate (Fe *trans* to R) **2**, which undergoes highly stereoselective alkylation reactions (Scheme 1). These enolate reactions require strong base (butyllithium) and low temperatures (-78°C). In contrast the derived methoxycarbene salt $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{OMe})(\text{CH}_2\text{R})\}]^+\text{BF}_4^-$ **3**^{4,5} may be deprotonated at ambient temperature with a mild base such as methoxide to generate the corresponding *Z*-enol ether (Fe *cis* to R) **4**. Alkylation of complex **4** is highly stereoselective and demethylation of the thus formed methoxycarbene cation with iodide generates the corresponding elaborated acyl complex (Scheme 1).⁶ The two sequences described above are stereocomplementary given the same face selectivity induced by the iron chiral auxiliary but opposite enolate to enol ether geometries.

We were interested in extending this chemistry to the nitrogen series to see if analogous stereoselective reactions could be achieved *via* α -metallated imines or enamines attached to the iron chiral auxiliary $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)]$. Achiral α -metallated imines attached to $[\text{Cr}(\text{CO})_5]$ have been reported by Hegedus *et al.*⁷ We describe here the synthesis and properties of a range of aminocarbene cations of the type $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{NHR}^2)(\text{CH}_2\text{R}^1)\}]^+$. Of particular relevance to the present study is the report by Davison and Reger⁸ of the reaction of racemic $(RS)-[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{Me})(\text{OEt})\}]^+\text{BF}_4^-$ with $(S)-(-)\alpha$ -methylbenzylamine yielding a separable mixture of diastereoisomers of $(R,S)-\dagger$ and $(S,S)-[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{Me})[\text{NHCH}(\text{Me})\text{Ph}]]^+\text{BF}_4^-$. Although the absolute configurations at iron for these two diastereoisomers could not be assigned Cotton effect studies were consistent with them being epimeric at iron as expected.

Results and Discussion

Methoxycarbene complexes $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{CH}_2\text{R})(\text{OMe})\}]^+\text{BF}_4^-$ may be readily prepared by treatment of $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\text{Br}]$ with the appropriate lithium acetylide followed by protonation to the corresponding vinylidene complexes^{9a} and subsequent addition of methanol.^{9b}

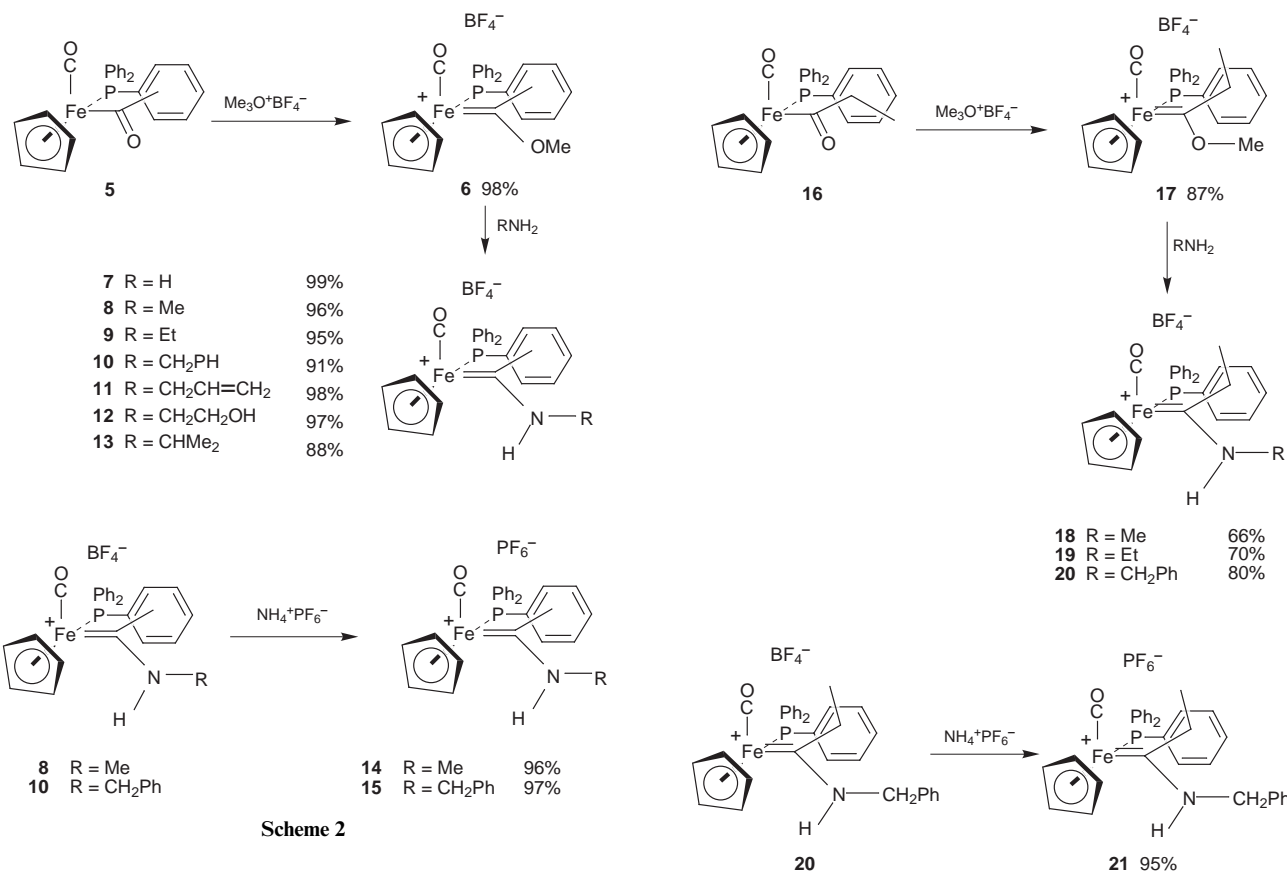
† The configuration at iron is given first followed by the configuration at the α -benzylic centre.



Scheme 1

Alternatively they may be synthesised from the corresponding acyl complexes $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)(\text{COCH}_2\text{R})]$ by treatment with trimethyloxonium tetrafluoroborate.¹⁰ Thus *O*-methylation of the iron acetyl complex $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)(\text{COMe})]$ **5** generated the methoxycarbene derivative $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{C}(\text{Me})(\text{OMe})\}]^+\text{BF}_4^-$ **6** in 98% yield (Scheme 2). Treatment of a tetrahydrofuran solution of the methoxycarbene complex **6** at 20°C with a large excess of ammonia gas or ten molar excess of a range of primary amines generated the corresponding aminocarbene complexes **7–13** in high yield (Scheme 2). If a large excess of amine was not employed then the reactions tended to be slower and the yields were lower due to *O*-demethylation of **6** back to **5**. Two of the tetrafluoroborate salts **8** and **10** were converted into their corresponding hexafluorophosphate salts **14** and **15** in excellent yield by treatment with aqueous $\text{NH}_4^+\text{PF}_6^-$ because of the better solubility of the latter salts in organic solvents.

Reaction of methoxycarbene complex **6** with aniline failed to give any of the corresponding aminocarbene complex, exclusively regenerating the acetyl complex **5** with concomitant formation of *N*-methylaniline. This change of reactivity may be



Scheme 2

attributed to reversible formation of the intermediate generated by addition of the amine to the methoxycarbene in the latter reaction with aniline but not in the former cases with primary aliphatic amines. Consistent with the observation of Davison and Reger,⁸ attempted coupling between secondary amines (Me₂NH, Et₂NH or pyrrolidine) and the methoxycarbene **6** proved to be very difficult with the product mixture consisting mostly of the demethylated neutral acetyl complex **5** and recovered methoxycarbene **6**. One aminocarbene complex derived from a secondary amine prepared *via* a different route has been reported.^{9b} Although some resonances in the ¹H NMR spectrum of the reaction mixture were consistent with those expected for the product aminocarbene complexes, these never exceeded 10% of the crude reaction mixture which could not be purified.

O-Methylation of the propanoyl complex **16** generated the methoxycarbene **17** in 87% yield, which reacted with selected primary amines to generate the corresponding aminocarbene complexes **18–20** (Scheme 3). Similar reactions of the pentanoyl complex **22** resulted, *via* the intermediacy of the methoxycarbene **23**, in the formation of the aminocarbenes **24** and **25** (Scheme 3).

For comparison with the known chemistry of the iron acyl and methoxycarbene complexes (Scheme 1), deprotonation–electrophilic trapping experiments were performed on complex **15**. The following deuteration experiments were readily monitored by ¹H NMR spectroscopy and by mass spectrometry. In deuteriomethanol solution containing a catalytic amount of methoxide complete H/D exchange was observed for the NH and CH₃ protons consistent with expected intermediacy under these equilibrating conditions of **26** and **27** (Scheme 4).

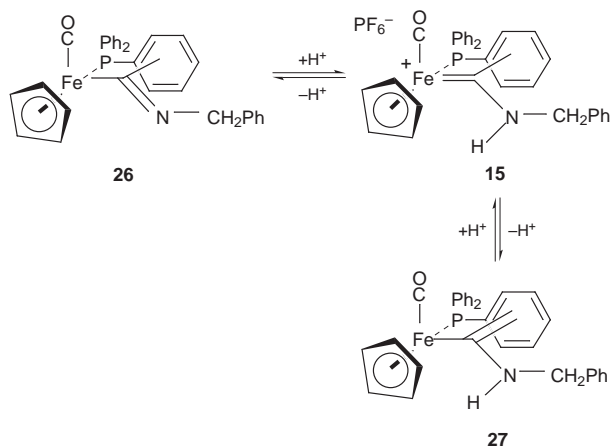
Under kinetically controlled deprotonation conditions treatment of **15** with 1 equivalent of butyllithium followed by quenching with deuteriomethanol led to the exchange of only the NH proton and formation of **28** (Scheme 5). In the presence of greater than 3 equivalents of butyllithium complex **15** generated, after quenching, exclusively the trideuterio derivative **30** (Scheme 5). The formation of **30** is consistent with the

Scheme 3

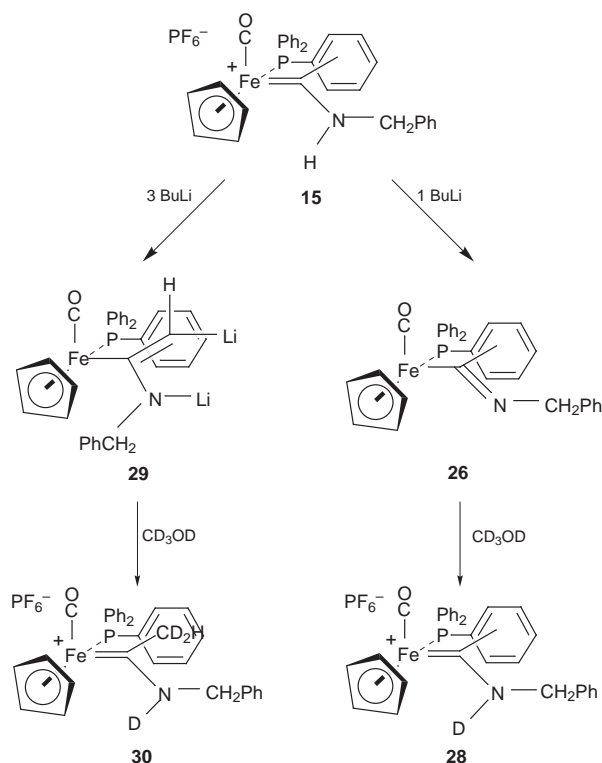
intermediacy of the dilithio species **29** for which there is some precedent in organic systems.¹¹

Treatment of **15** with 2 equivalents of butyllithium followed by quenching with deuteriomethanol led to a mixture of **28**, **30** and **32** rather than exclusive formation of the monolithio species **31** and hence after quenching to **32**. All attempts to isolate complex **26** were unsuccessful, only complex **15** being recovered in each case. In all of the above experiments attempted quenchings with methyl iodide were completely unsuccessful indicating that the intermediates **26**, **29** and **31** are unreactive towards alkylation. This observation is in direct contrast to the reactions of the enolates **2** and enol ethers **4**.

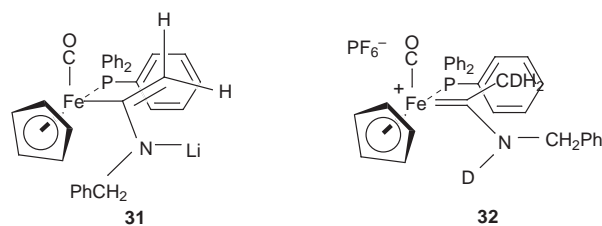
Very few methods exist for the resolution of the iron chiral auxiliary [(η⁵-C₅H₅)Fe(CO)(PPh₃)]¹² and therefore we were interested to see if any chiral recognition could be achieved



Scheme 4

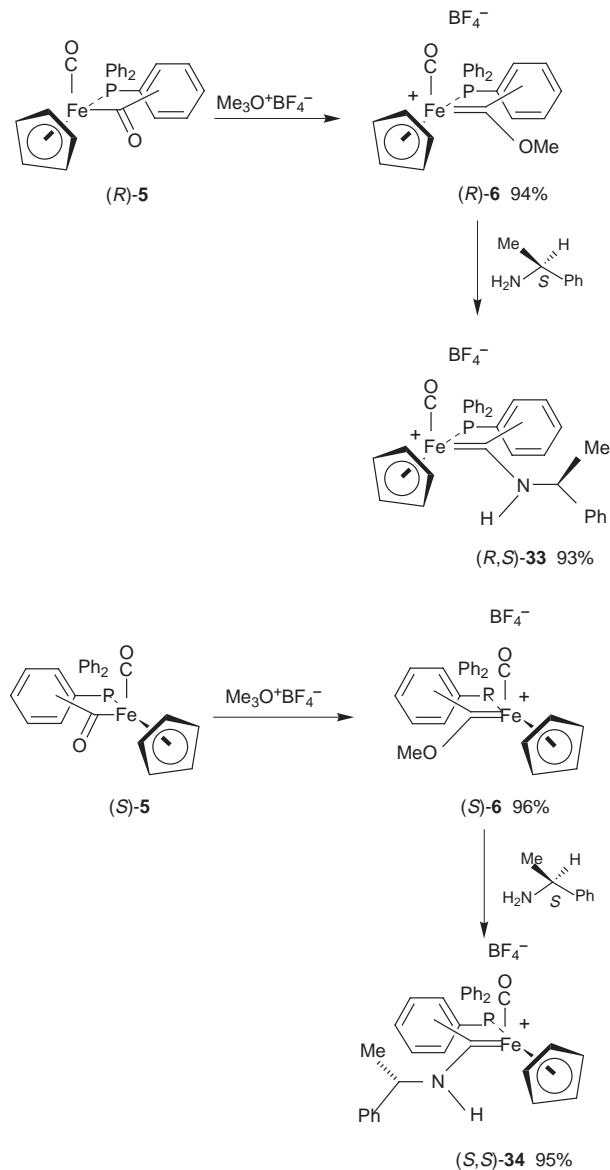


Scheme 5



during the reaction of the methoxycarbene complex **6** and α -methylbenzylamine. Significant chiral recognition would allow the kinetic resolution of racemic **6** with homochiral α -methylbenzylamine. *O*-Methylation of homochiral (*R*)-**5** generated the homochiral methoxycarbene complex (*R*)-**6** (Scheme 6). Treatment of (*R*)-**6** with (*S*)- α -methylbenzylamine generated the corresponding aminocarbene complex (*R,S*)-**33** as a single diastereoisomer. Similar treatment of (*S*)-**5** gave *via* (*S*)-**6**¹³ the aminocarbene complex epimeric at iron (*S,S*)-**34**.

The diastereoisomers (*R,S*)-**33** and (*S,S*)-**34** were readily distinguishable by NMR spectroscopy and their configurational assignments were confirmed by an X-ray crystal-structure analysis of (*R,S*)-**33** which confirmed retention of configura-



Scheme 6

Table 1 Selected bond distances (Å) and torsion angles (°) for (*R,S*)-**33**

Fe(1)–C(1)	1.928(6)	C(1)–N(1)	1.318(7)
C(1)–C(2)	1.513(8)	N(1)–C(3)	1.495(6)
C(1)–N(1)–C(3)–C(4)	–124	C(1)–N(1)–C(3)–C(10)	110
C(2)–C(1)–N(1)–C(3)	–6	Fe(1)–C(1)–N(1)–C(3)	177
N(1)–C(3)–C(4)–C(5)	127		

tion at both the iron and the α -methylbenzylamine stereogenic centres during the conversion of the methoxycarbene to the aminocarbene. Fig. 1 shows the molecular conformation of (*R,S*)-**33** present in the crystal structure. Selected bond lengths and torsion angles are given in Table 1. The structure of (*R,S*)-**33** shows the normal pseudo-octahedral geometry around iron with the nitrogen of the aminocarbene *anti* to the carbon monoxide ligand. The α -methylbenzyl group is antiperiplanar to the iron auxiliary with the benzylic hydrogen synperiplanar to the aminocarbene methyl group. The conformation adopted in the solid state parallels that of the corresponding methoxycarbene complex¹⁴ and is as expected on steric grounds.

In order to investigate the possible chiral recognition phenomenon, racemic methoxycarbene complex **6** was treated with racemic α -methylbenzylamine in a variety of solvents. Employing all racemic conditions allows a direct correlation of the

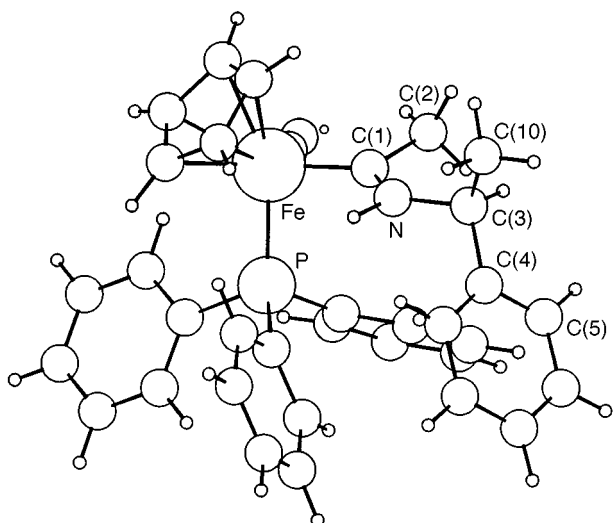


Fig. 1 Crystal structure of (R,S) - $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Me})\text{-NHCH}(\text{Me})\text{Ph}\}]^+\text{BF}_4^-$ (R,S)-**33**

ratio of product diastereoisomers with the stereoselectivity factor without the need to consider the complicating effects of mass action. Unfortunately, in a range of solvents (tetrahydrofuran, methanol, ethanol, propan-2-ol, hexane, acetone, ethyl acetate, dibutyl ether, benzene, chloroform and carbon tetrachloride) equimolar mixtures of product diastereoisomers were observed within experimental error. The complete lack of chiral recognition does not allow a kinetic resolution procedure to be effected. However, Davison and Reger⁸ have demonstrated the classical resolution procedure by separation of (R,S) -**33** and (S,S) -**34** by fractional crystallisation. Photochemical studies on the above aminocarbene complexes have been reported elsewhere.¹⁵

Experimental

All reactions and purifications were performed under a nitrogen atmosphere using standard vacuum line and Schlenk-tube techniques.¹⁶ All solvents were deoxygenated before use. Tetrahydrofuran (THF) and dichloromethane were distilled from sodium benzophenone and calcium hydride, respectively, under nitrogen. Analytical grade acetone (FSA) was used without purification. Unless otherwise indicated, all commercially available reagents were used as received. Column chromatography was performed on grade I (activated) basic alumina or on silica gel (Merck Kieselgel 60). Elemental analyses were carried out by the Dyson Perrins Laboratory Analytical Service. Butyllithium (1.5 M in hexane) and methyllithium (1.4 M in diethyl ether) were used as supplied by Aldrich. Infrared spectra were recorded on either a Perkin-Elmer 781 or a Perkin-Elmer 1750 Fourier-transform spectrophotometer in dichloromethane solutions using 1 mm NaCl cells unless otherwise stated. Proton NMR spectra were recorded on either a Varian-Gemini 200 (200 MHz) or a Bruker AM500 (500.13 MHz) spectrometer in CDCl_3 solutions unless otherwise stated. Carbon-13, ^{19}F and ^{31}P NMR spectra were recorded on a Bruker AM250 spectrometer (at 62.90, 235.35 and 125.76 MHz respectively) in CDCl_3 solutions. Proton and ^{13}C NMR spectra were referenced to tetramethylsilane using internal solvent peaks. In the Experimental section, ^{13}C , ^{19}F and ^{31}P NMR data are described in terms of the proton decoupled (broad band) spectra. Abbreviations used: br, broad single; d, doublet; m, multiplet; q, quartet; s, singlet; t, triplet. Mass spectra were recorded on VG Micromass ZAB1F or MM30F instruments using FAB techniques for organometallic compounds. Optical rotations were recorded on a Perkin-Elmer 241 polarimeter in the solvents indicated at 20 °C in a 1 dm cell. Concentrations (c) are given in $\text{g } 100 \text{ mL}^{-1}$ and specific rotation in units of 10^{-1} deg

$\text{cm}^2 \text{ g}^{-1}$. The complexes (RS) - $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)(\text{COMe})]$ (RS)-**5**,¹⁷ (R)-**5**,¹² (S)-**5**,¹² (RS)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)(\text{COEt})]$ (RS)-**16**,¹⁸ (RS)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)(\text{COBu})]$ (RS)-**22**,^{6,19} (RS)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Me})(\text{NHMe})\}]^+\text{BF}_4^-$ (RS)-**8**,²⁰ (RS)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Bu})(\text{OMe})\}]^+\text{BF}_4^-$ (RS)-**23**²⁰ and (RS)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Bu})(\text{NHCH}_2\text{-Ph})\}]^+\text{BF}_4^-$ (RS)-**25**²⁰ were synthesised by literature procedures. (S)-(-)- α -Methylbenzylamine was purchased from Aldrich [α]_D (neat) -39.0, this corresponds to an enantiomeric excess of 95% {lit.,^{21a} [α]_D (neat) -40.6}. However with the method of Parker and Taylor^{21b} using ^1H NMR with (R)-(-)- O -acetylmandelic acid [$\text{MeCO}_2\text{CH}(\text{C}_6\text{H}_5)\text{CO}_2\text{H}$] an enantiomeric excess of 99.6% was obtained.

General procedure for the synthesis of the methoxycarbenes

A nitrogen degassed solution of $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)(\text{COR})]$ ($R = \text{Me, Et or Bu}$) (30 mL per 1 g of complex) was cannulated into a flask containing trimethyloxonium tetrafluoroborate (1.3 mol equivalents, dried at 60 °C *in vacuo* for 3 h) and the resultant mixture stirred under a nitrogen atmosphere for 8–21 h. During this time the colour of the reaction mixture changed from a deep orange-red to a deep yellow-brown. The reaction mixture was then filtered *via* cannula and concentrated *in vacuo* to yield a yellow-brown oil which was dissolved in a minimum amount of dichloromethane and cannulated dropwise into diethyl ether (40 mL g^{-1}) cooled to 0 °C, forming a bright yellow precipitate. The solvent was removed *via* cannula and the residue washed twice with diethyl ether before being dried *in vacuo*. For analytical analysis the methoxycarbenes were recrystallised from the solvents indicated.

(RS)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Me})(\text{OMe})\}]^+\text{BF}_4^-$ (RS)-6**.** Methylation of (R,S)-**5** (20.0 g, 44.1 mmol) with $\text{Me}_3\text{O}^+\text{BF}_4^-$ (8.46 g, 57.2 mmol) for 8 h at room temperature afforded (RS)-**6** as a yellow solid (23.99 g, 98%) which was crystallised from $\text{CH}_2\text{Cl}_2\text{-Et}_2\text{O}$ as yellow microcrystals, m.p. 162–165 °C (decomp.) (lit.,^{10a} 164 °C); ν_{max} 1966 (Fe–CO) and 1057 cm^{-1} (BF_4^-); δ_{H} 2.84 (3 H, s, Fe=CMe), 3.99 (3 H, s, Fe=COMe), 4.92 (5 H, d, J_{PH} 1.2 Hz, C_5H_5), 7.20–7.40 (6 H, m, H_o of PPh_3), 7.43–7.65 (9 H, m, H_m and H_p of PPh_3); δ_{C} 44.70 (s, Fe=CMe), 65.24 (s, Fe=COMe), 87.75 (s, C_5H_5), 129.12 (d, J_{PC} 11.0, C_m of PPh_3), 131.16 (s, C_p of PPh_3), 132.54 (d, J_{PC} 9.5, C_o of PPh_3), 132.72 (d, J_{PC} 47.4, C_{ipso} of PPh_3), 215.47 (d, J_{PC} 27.4, Fe–CO), 263.19 (d, J_{PC} 21.4 Hz, Fe=C); δ_{P} 63.13; δ_{F} -154.49 and -154.54 (1 : 3); m/z 469 (M^+ , 87), 441 ($M - \text{CO}$, 86), 439 ($M - 2\text{Me}$, 10), 383 (439 - 2CO, 100), 318 (12), 295 (8), 263 (24), 239 (17), 183 (55), 164 (10), 108 (8), 85 (8%).

(R)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Me})(\text{OMe})\}]^+\text{BF}_4^-$ (R)-6**.** Methylation of (R)-**5**¹² (1.42 g, 3.13 mmol) with $\text{Me}_3\text{O}^+\text{BF}_4^-$ (0.60 g, 4.06 mmol) for 12 h at room temperature afforded (R)-**6** as a yellow solid (1.64 g, 94%) which was crystallised from $\text{CH}_2\text{Cl}_2\text{-Et}_2\text{O}$ as yellow microcrystals, m.p. 162–164 °C (decomp.) (Found: C, 58.19; H, 4.45. $\text{C}_{27}\text{H}_{26}\text{BF}_4\text{FeO}_2\text{P}$ requires: C, 58.31; H, 4.72%); [α]_D (20 °C, $c = 3.2$, CH_2Cl_2) -257.0.

(S)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Me})(\text{OMe})\}]^+\text{BF}_4^-$ (S)-6**.** Methylation of (S)-**5**¹² (2.00 g, 4.41 mmol) with $\text{Me}_3\text{O}^+\text{BF}_4^-$ (0.85 g, 5.73 mmol) for 12 h at room temperature afforded (S)-**6**¹¹ as a yellow solid (2.35 g, 96%) which was crystallised from $\text{CH}_2\text{Cl}_2\text{-Et}_2\text{O}$ as yellow microcrystals, m.p. 162–165 °C (decomp.) (Found: C, 58.07; H, 4.38. $\text{C}_{27}\text{H}_{26}\text{BF}_4\text{FeO}_2\text{P}$ requires: C, 58.31; H, 4.72%); [α]_D (20 °C, $c = 2.3$, CH_2Cl_2) 257.7.

(RS)- $[(\eta^5\text{-C}_5\text{H}_5)\text{Fe}(\text{CO})(\text{PPh}_3)\{\text{=C}(\text{Et})(\text{OMe})\}]^+\text{BF}_4^-$ (RS)-17**.** Methylation of (RS)-**16** (5.0 g, 10.7 mmol) with $\text{Me}_3\text{O}^+\text{BF}_4^-$ (2.04 g, 13.9 mmol) for 21 h at room temperature afforded (RS)-**17** as a yellow solid (5.29 g, 87%) which was

crystallised from $\text{CH}_2\text{Cl}_2\text{-Et}_2\text{O}$ as yellow microcrystals, m.p. 169–172 °C (decomp.) (Found: C, 58.70; H, 4.86. $\text{C}_{28}\text{H}_{28}\text{BF}_4\text{FeO}_2\text{P}$ requires: C, 58.98; H, 4.95%; ν_{max} 1966 (Fe–CO) and 1052 cm^{-1} (BF_4^-); δ_{H} 1.16 (3 H, d, J 7.6, Fe=CCH₂Me), 2.76 (2 H, q, J 7.6, Fe=CCH₂Me), 4.25 (3 H, s, Fe=COMe), 4.89 (5 H, d, J_{PH} 1.4 Hz, C₅H₅), 7.20–7.40 (6 H, m, H_o of PPh₃), 7.45–7.60 (9 H, m, H_m and H_p of PPh₃); δ_{C} 10.17 (s, Fe=CCH₂Me), 49.75 (s, Fe=CCH₂Me), 65.84 (s, Fe=COMe), 87.31 (s, C₅H₅), 129.15 (d, J_{PC} 9.9, C_m of PPh₃), 131.41 (s, C_p of PPh₃), 132.75 (d, J_{PC} 10.4, C_o of PPh₃), 132.84 (d, J_{PC} 47.3, C_{ipso} of PPh₃), 216.26 (d, J_{PC} 26.8, Fe–CO), 257.91 (d, J_{PC} 21.1 Hz, Fe=C); δ_{F} –154.69 and –154.75 (1:3); m/z 483 (M^+ , 100), 455 (M – CO, 86), 439 (M – Me – Et, 14), 383 (M – Me – Et, 100), 318 (18), 295 (38), 263 (36), 239 (19), 221 (16), 193 (38), 183 (65), 161 (24), 121 (31), 103 (12), 85 (25%).

General procedure for the synthesis of the aminocarbenes

The amine (10 mol equivalents) was injected dropwise into a solution of the methoxycarbene in THF (20 mL g^{-1}) under a nitrogen atmosphere and the resultant red-orange solution stirred at the specified temperatures for the times indicated. The reaction mixture was then concentrated *in vacuo* to yield a red oil which was dissolved in a minimum amount of dichloromethane and cannulated dropwise into diethyl ether (40 mL g^{-1}) cooled to 0 °C forming a bright yellow precipitate. The solvent was removed *via* cannula and the residue washed twice with diethyl ether before being dried *in vacuo*. For analytical analysis the aminocarbenes were recrystallised from the solvents indicated.

(RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₃]=C(Me)(NH₂)]⁺BF₄⁻ (RS)-7. Reaction of (RS)-6 (0.91 g, 1.64 mmol) with gaseous NH₃ for 1.5 h at room temperature afforded (RS)-7 as a yellow solid (0.87 g, 99%) which was crystallised from $\text{CH}_2\text{Cl}_2\text{-Et}_2\text{O}$ as yellow microneedles, m.p. 93–97 °C (decomp.) (Found: C, 57.89; H, 4.85; N, 2.62; P, 5.57. $\text{C}_{26}\text{H}_{25}\text{BF}_4\text{FeNOP}\cdot 0.67\text{C}_2\text{H}_{10}\text{O}$ requires: C, 57.87; H, 4.86; N, 2.53; P, 5.60%; ν_{max} 1958 (Fe–CO) and 1098 cm^{-1} (BF_4^-); δ_{H} 2.40 (3 H, s, Fe=CMe), 4.76 (5 H, d, J_{PH} 1.2 Hz, C₅H₅), 7.30–7.33 (6 H, m, H_o of PPh₃), 7.50–7.53 (9 H, m, H_m and H_p of PPh₃), 9.26 (1 H, br s, NH₂), 9.37 (1 H, br s, NH₂); δ_{C} 43.75 (s, Fe=CMe), 85.65 (s, C₅H₅), 129.16 (d, J_{PC} 10.2, C_m of PPh₃), 131.17 (s, C_p of PPh₃), 132.78 (d, J_{PC} 9.4, C_o of PPh₃), 133.03 (d, J_{PC} 44.7, C_{ipso} of PPh₃), 217.38 (d, J_{PC} 28.1, Fe–CO), 273.89 (d, J_{PC} 19.4 Hz, Fe=C); δ_{P} 66.46; δ_{F} –153.06 and –153.11 (1:3); m/z 454 (M^+ , 100), 426 (M – CO, 58), 383 (24), 295 (30), 263 (42), 183 (35), 164 (82), 121 (18), 85 (18%).

(RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₃]=C(Me)(NH₂Et)]⁺BF₄⁻ (RS)-9. Reaction of (RS)-6 (1.00 g, 1.80 mmol) with gaseous ethylamine for 2 h at room temperature afforded (RS)-9 as a yellow solid (0.97 g, 95%) which was crystallised from $\text{CH}_2\text{Cl}_2\text{-Et}_2\text{O}$ as red-brown-yellow microcrystals, m.p. 184–187 °C (decomp.) (Found: C, 58.86; H, 5.09; N, 2.34; P, 5.57. $\text{C}_{28}\text{H}_{29}\text{BF}_4\text{FeNOP}$ requires: C, 59.09; H, 5.14; N, 2.46; P, 5.44%; ν_{max} 1954 (Fe–CO) and 1056 cm^{-1} (BF_4^-); δ_{H} 0.93 (3 H, t, J 7.1, NCH₂CH₃), 2.39 (3 H, s, Fe=CMe), 3.45 (2 H, m, NCH₂CH₃), 4.77 (5 H, d, J_{PH} 1.1 Hz, C₅H₅), 7.20–7.30 (6 H, m, H_o of PPh₃), 7.40–7.50 (9 H, m, H_m and H_p of PPh₃), 9.57 (1 H, br s, NH₂Et); δ_{C} 13.20 (s, NCH₂CH₃), 36.92 (s, Fe=CMe), 44.88 (s, NCH₂CH₃), 85.33 (s, C₅H₅), 129.05 (d, J_{PC} 9.8, C_m of PPh₃), 131.01 (s, C_p of PPh₃), 132.69 (d, J_{PC} 9.2, C_o of PPh₃), 133.19 (d, J_{PC} 45.1, C_{ipso} of PPh₃), 217.66 (d, J_{PC} 28.8, Fe–CO), 263.69 (d, J_{PC} 21.0 Hz, Fe=C); δ_{P} 67.65; δ_{F} –153.80 and –153.85 (1:3); m/z 482 (M^+ , 73), 454 (M – CO, 28), 383 (9), 318 (10), 295 (8), 263 (14), 192 (100%).

(RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₃]=C(Me)(NHMe)]⁺PF₆⁻ (RS)-14. A saturated solution of NH₄PF₆ in water (150 mL) was added dropwise to a solution of (RS)-8 (1.01 g, 1.78 mmol)

in acetone (40 mL) at room temperature under a nitrogen atmosphere. This resulted in the immediate precipitation of a bright yellow solid which was filtered off and washed with water, ethanol and diethyl ether to afford pure (RS)-14 (1.07 g, 96%) which was crystallised from $\text{CH}_2\text{Cl}_2\text{-Et}_2\text{O}$ as red-brown yellow-brown plates, m.p. 172–174 °C (decomp.) (Found: C, 53.49; H, 4.38; N, 2.20. $\text{C}_{28}\text{H}_{29}\text{F}_6\text{FeNOP}_2$ requires: C, 53.61; H, 4.66; N, 2.23%; ν_{max} 1957 (Fe–CO) and 847 cm^{-1} (PF_6^-); δ_{H} (CD₃COCD₃) 1.04 (3 H, t, J 7.3, NCH₂CH₃), 2.62 (3 H, s, Fe=CMe), 3.52 (1 H, dq, J 13.2, 7.0, NCH₂CH₃), 3.59 (1 H, dq, J 13.3, 7.1, NCH₂CH₃), 4.98 (5 H, s, C₅H₅), 7.38 (6 H, dd, J_{PH} 10.3, J_{HH} 7.3 Hz, H_o of PPh₃), 7.56–7.63 (9 H, m, H_m and H_p of PPh₃), 9.79 (1 H, br s, NH₂Et); δ_{C} (CD₃COCD₃) 13.60 (s, NCH₂CH₃), 37.52 (s, Fe=CMe), 49.57 (s, NCH₂CH₃), 86.46 (s, C₅H₅), 129.99 (d, J_{PC} 9.9, C_m of PPh₃), 131.94 (s, C_p of PPh₃), 133.79 (d, J_{PC} 9.9, C_o of PPh₃), 133.37 (d, J_{PC} 45.5, C_{ipso} of PPh₃), 218.67 (d, J_{PC} 29.3, Fe–CO), 266.67 (d, J_{PC} 22.1 Hz, Fe=C); δ_{P} (CD₃COCD₃) 73.12, –142.92 (septet, J_{PF} 708.5 Hz, PF_6^-); δ_{F} (CD₃COCD₃) –67.48 (d, J_{PF} 708.5 Hz, PF_6^-); m/z 482 (M^+ , 65), 454 (M – CO, 34), 383 (10), 318 (10), 263 (14), 192 (100), 183 (20), 121 (18), 70 (20%).

(RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₃]=C(Me)(NHCH₂Ph)]⁺BF₄⁻ (RS)-10. Reaction of (RS)-6 (24.27 g, 44.0 mmol) with benzylamine (42.6 mL, 0.43 mol) at 0 °C for 2 h afforded (RS)-10 as a yellow solid (25.12 g, 91%). Recrystallisation from $\text{CH}_2\text{Cl}_2\text{-EtOH}$ gave yellow microneedles, m.p. 199–201 °C (decomp.) [lit.,^{9b} 203 °C (decomp.)] (Found: C, 63.08; H, 4.92; N, 2.23; P, 4.96. $\text{C}_{33}\text{H}_{31}\text{BF}_4\text{FeNOP}$ requires: C, 62.79; H, 4.95; N, 2.22; P, 4.91%; ν_{max} 1960 (Fe–CO) and 1062 cm^{-1} (BF_4^-); δ_{H} (CDCl₃) 2.40 (3 H, s, Fe=CMe), 4.66 (1 H, dd, J 14.4, 6.8, NCH₂Ph), 4.76 (1 H, dd, J 14.4, 5.5, NCH₂Ph), 4.90 (5 H, d, J_{PH} 1.6, C₅H₅), 7.02 (2 H, br d, J 7.0, H_o of Ph), 7.11 (6 H, dd, J_{PH} 10.9, J_{HH} 7.7, H_o of PPh₃), 7.20–7.29 (3 H, m, H_m and H_p of Ph), 7.40 (6 H, td, J 7.4, 2.0, H_m of PPh₃), 7.48 (3 H, td, J 7.9, 2.0 Hz, H_p of PPh₃), 10.14 (1 H, br s, NHCH₂Ph); δ_{C} (CD₃COCD₃) 2.67 (3 H, s, Fe=CMe), 4.74 (1 H, d, J 13.2, NCH₂Ph), 4.82 (1 H, d, J 13.3, NCH₂Ph), 5.02 (5 H, d, J_{PH} 0.9, C₅H₅), 7.20 (2 H, dd, J 7.5, 2.1, H_o of Ph), 7.30 (6 H, dd, J_{PH} 10.7, J_{HH} 7.8, H_o of PPh₃), 7.35–7.40 (3 H, m, H_m and H_p of Ph), 7.50 (6 H, td, J 7.7, 2.0, H_m of PPh₃), 7.58 (3 H, td, J 7.2, 1.8 Hz, H_p of PPh₃), 10.32 (1 H, br s, NHCH₂Ph); δ_{C} (CD₃COCD₃) 37.47 (s, Fe=CMe), 52.49 (s, NCH₂Ph), 85.57 (s, C₅H₅), 127.69 (s, C_p of Ph), 127.89 (s, C_m of Ph), 128.76 (s, C_o of Ph), 128.91 (d, J_{PC} 9.2, C_m of PPh₃), 130.83 (s, C_p of PPh₃), 132.53 (d, J_{PC} 9.7, C_o of PPh₃), 133.25 (d, J_{PC} 44.4, C_{ipso} of PPh₃), 135.45 (s, C_{ipso} of Ph), 217.85 (d, J_{PC} 28.8, Fe–CO), 265.44 (d, J_{PC} 23.9 Hz, Fe=C); δ_{P} (CD₃COCD₃) 67.31; δ_{F} (CD₃COCD₃) –153.42 and –153.48 (1:3); m/z 544 (M^+ , 37), 516 (M – CO, 7), 383 (7), 295 (9), 254 (38), 152 (47), 135 (42), 103 (34), 85 (95), 59 (70), 47 (100%).

(RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₃]=C(Me)(NHCH₂Ph)]⁺PF₆⁻ (RS)-15. A saturated solution of NH₄PF₆ in water (150 mL) was added dropwise to a solution of (RS)-10 (1.00 g, 1.58 mmol) in acetone (200 mL) at room temperature under a nitrogen atmosphere. This resulted in the immediate precipitation of a bright yellow solid which was filtered off, washed with water, ethanol and diethyl ether to afford pure (RS)-15 (1.06 g, 97%). Recrystallisation from $\text{CH}_2\text{Cl}_2\text{-acetone}$ gave yellow microneedles, m.p. 190–192 °C (Found: C, 57.34; H, 4.32; N, 2.06. $\text{C}_{33}\text{H}_{31}\text{F}_6\text{FeNOP}_2$ requires: C, 57.49; H, 4.54; N, 2.03%; ν_{max} 1959 (Fe–CO) and 841 cm^{-1} (PF_6^-); δ_{H} (CD₃COCD₃) 2.69 (3 H, d, J 3.0, Fe=CMe), 4.73 (1 H, dd, J 14.8, 3.8, NHCH₂Ph), 4.83 (1 H, dd, J 14.8, 5.3, NHCH₂Ph), 5.02 (5 H, d, J_{PH} 1.5, C₅H₅), 7.20 (2 H, dd, J 7.6, 2.0, H_o of Ph), 7.30 (6 H, dd, J_{PH} 10.8, J_{HH} 7.7, H_o of PPh₃), 7.35–7.42 (3 H, m, H_m and H_p of Ph), 7.50 (6 H, td, J 7.9, 2.3, H_m of PPh₃), 7.60 (3 H, tq, J 7.8, 1.9 Hz, H_p of PPh₃), 10.26 (1 H, br s, NHCH₂Ph); δ_{C} (CD₃COCD₃) 38.29 (s, Fe=CMe), 54.14 (s, NCH₂Ph), 86.36 (s, C₅H₅), 129.00 (s, C_p of Ph), 129.19 (s, C_m of Ph), 129.88 (s, C_o of Ph), 129.91 (d, J_{PC}

8.3, C_m of PPh_3), 131.90 (s, C_p of PPh_3), 133.50 (d, J_{PC} 9.6, C_o of PPh_3), 133.67 (d, J_{PC} 31.4, C_{ipso} of PPh_3), 134.17 (s, C_{ipso} of Ph), 218.85 (d, J_{PC} 30.8, Fe–CO), 271.44 (d, J_{PC} 23.1 Hz, Fe=C); $\delta_P(CD_3COCD_3)$ 67.49 (s, PPh_3), –147.12 (septet, J_{PF} 706.5 Hz, PF_6^-); $\delta_F(CD_3COCD_3)$ –69.18 (d, J_{PF} 707.7 Hz, PF_6^-); m/z 544 (M^+ , 85), 516 (M – CO, 13), 383 (14), 279 (19), 254 (100), 212 (14), 183 (20), 91 (20), 85 (24), 59 (18), 47 (27%).

(*RS*)-[(η^5 - C_5H_5)Fe(CO)(PPh₃)₂]=C(Me)(NHCH₂CH=CH₂)]⁺BF₄[–] (RS-11**). Reaction of (*RS*)-6 (0.50 g, 0.90 mmol) with allylamine (1.18 g, 20 mmol) at room temperature for 2 h afforded (*RS*)-11 as a yellow solid (0.51 g, 98%). Recrystallisation from CH₂Cl₂–EtO₂ gave red microneedles, m.p. 174–177 °C (Found: C, 59.77; H, 4.86; N, 1.97; P, 5.18. C₂₉H₂₉BF₄FeNOP requires: C, 59.93; H, 5.03; N, 2.41; P, 5.33%); ν_{max} 1955 (Fe–CO) and 1055 cm^{–1} (BF₄[–]); δ_H 2.37 (3 H, s, Fe=CMe), 4.00–4.10 (2 H, m, NCH₂CH=CH₂), 4.78 (5 H, s, C₅H₅), 5.08 (1 H, d, J 12.5, NCH₂CH=CH₂), 5.11 (1 H, d, J 4.5 Hz, NCH₂CH=CH₂), 5.43–5.52 (1 H, m, NCH₂CH=CH₂), 7.24–7.27 (6 H, m, H_o of PPh₃), 7.48–7.50 (9 H, m, H_m and H_p of PPh₃), 9.78 (1 H, br s, NHCH₂CH=CH₂); δ_C 37.46 (s, Fe=CMe), 52.44 (s, NCH₂CH=CH₂), 85.58 (s, C₅H₅), 119.95 (s, NCH₂CH=CH₂), 129.14 (d, J_{PC} 9.5, C_m of PPh₃), 130.67 (s, NCH₂CH=CH₂), 131.09 (s, C_p of PPh₃), 132.80 (d, J_{PC} 9.5, C_o of PPh₃), 133.26 (d, J_{PC} 45.7, C_{ipso} of PPh₃), 217.80 (d, J_{PC} 29.0, Fe–CO), 265.96 (d, J_{PC} 22.6 Hz, Fe=C); δ_P 67.05; δ_F –153.77 and –153.83 (1:3); m/z 494 (M^+ , 100), 466 (M – CO, 22), 458 (9), 383 (20), 310 (12), 295 (18), 263 (30), 239 (11), 227 (17), 204 (87), 183 (37), 162 (20), 121 (18%).**

(*RS*)-[(η^5 - C_5H_5)Fe(CO)(PPh₃)₂]=C(Me)(NHCH₂CH₂OH)]⁺BF₄[–] (RS-12**). Reaction of (*RS*)-6 (1.80 g, 3.24 mmol) with aminoethanol (1.95 mL, 32.4 mmol) at room temperature for 15 min afforded (*RS*)-12 as an orange solid (2.08 g, 97%). Recrystallisation from CH₂Cl₂–EtO₂ gave orange microneedles, m.p. 139–141 °C (Found: C, 59.64; H, 5.18; N, 2.23; P, 5.05. C₂₈H₂₉BF₄FeNO₂P requires: C, 59.52; H, 5.00; N, 2.39; P, 5.29%); ν_{max} 1940 (Fe–CO) and 1057 cm^{–1} (BF₄[–]); δ_H 2.46 (3 H, s, Fe=CMe), 2.81 (1 H, t, J 5.5, NCH₂CH₂OH), 3.42–3.58 (3 H, m, NCH₂CH₂OH), 4.77 (5 H, d, J_{PH} 1.0, C₅H₅), 7.28 (6 H, dd, J_{PH} 10.9, J_{HH} 7.5, H_o of PPh₃), 7.50 (6 H, td, J 6.9, 1.9, H_m of PPh₃), 7.55 (3 H, td, J 7.5, 2.0 Hz, H_p of PPh₃), 9.35 (1 H, br s, NHCH₂CH₂OH); δ_C 38.11 (s, Fe=CMe), 52.10 (s, NCH₂CH₂OH), 59.83 (s, NCH₂CH₂OH), 85.97 (s, C₅H₅), 129.68 (d, J_{PC} 9.6 Hz, C_m of PPh₃), 131.70 (s, C_p of PPh₃), 133.37 (d, J_{PC} 10.0, C_o of PPh₃), 133.74 (d, J_{PC} 53.9, C_{ipso} of PPh₃), 218.25 (d, J_{PC} 28.4, Fe–CO), 267.64 (d, J_{PC} 22.1 Hz, Fe=C); δ_P 69.52; δ_F –153.54 and –153.58 (1:3); m/z 498 (M^+ , 71), 470 (M – CO, 8), 383 (13), 318 (6), 295 (35), 263 (25), 208 (100), 183 (22), 142 (13%).**

(*RS*)-[(η^5 - C_5H_5)Fe(CO)(PPh₃)₂]=C(Me)(NHCHMe₂)]⁺BF₄[–] (RS-13**). Reaction of (*RS*)-6 (1.11 g, 2.00 mmol) with isopropylamine (1.18 g, 20 mmol) at room temperature for 2 h afforded (*RS*)-13 as a yellow solid (1.02 g, 88%). Recrystallisation from CH₂Cl₂–EtOH gave yellow microcrystals, m.p. 213–214 °C (decomp.) (Found: C, 59.81; H, 5.58; N, 2.29; P, 5.15. C₂₉H₃₁BF₄FeNOP requires: C, 59.73; H, 5.35; N, 2.40; P, 5.31%); ν_{max} 1964 (Fe–CO) and 1058 cm^{–1} (BF₄[–]); δ_H 0.97 (3 H, d, J 6.5, CHMe₂), 1.23 (3 H, d, J 6.5, CHMe₂), 2.53 (3 H, s, Fe=CMe), 3.96 (1 H, octet, NHCHMe₂), 4.80 (5 H, s, C₅H₅), 7.19–7.27 (6 H, m, H_o of PPh₃), 7.48–7.53 (9 H, m, H_m and H_p of PPh₃), 9.22 (1 H, br s, NHCHMe₂); δ_C 20.67 (s, NHCHMe₂), 21.21 (s, NHCHMe₂), 37.38 (s, Fe=CMe), 52.51 (s, NHCHMe₂), 85.40 (s, C₅H₅), 129.25 (d, J_{PC} 9.8, C_m of PPh₃), 131.18 (s, C_p of PPh₃), 132.74 (d, J_{PC} 9.7, C_o of PPh₃), 133.07 (d, J_{PC} 44.8 Hz, C_{ipso} of PPh₃), 217.54 (d, J_{PC} 28.8, Fe–CO), 262.48 (d, J_{PC} 21.8 Hz, Fe=C); δ_P 68.32; δ_F –152.73 and –152.79 (1:3); m/z 496 (M^+ , 35), 468 (M – CO, 15), 383 (6), 318 (10), 294 (9), 206 (100), 183 (27), 164 (20), 121 (21), 84 (25), 56 (9%).**

(*R,S*)-[(η^5 - C_5H_5)Fe(CO)(PPh₃)₂]=C(Me)(NHCH(Me)Ph)]⁺BF₄[–] (R,S**-33). Reaction of (*R*)-6 (0.40 g, 0.72 mmol) with (*S*)- α -methylbenzylamine (0.93 mL, 7.19 mmol) at room temperature for 2 h afforded (*R,S*)-33 as a yellow solid (0.43 g, 93%). Recrystallisation from CH₂Cl₂–EtOH gave orange micro-needles, m.p. 194–196 °C (decomp.) [lit.,⁸ 195 °C (decomp.)] (Found: C, 63.51; H, 5.16; N, 2.16; P, 5.00. C₃₄H₃₃BF₄FeNOP requires: C, 63.29; H, 5.15; N, 2.17; P, 4.80%); $[\alpha]_D$ (20 °C, c = 3.00, CH₂Cl₂) –16.3; ν_{max} 1965 (Fe–CO) and 1058 cm^{–1} (BF₄[–]); $\delta_H(CD_3COCD_3)$ 1.60 [3 H, d, J 6.8 Hz, NHCH(Me)Ph], 2.64 (3 H, Fe=CMe), 5.01 (5 H, d, J_{PH} 1.3, C₅H₅), 5.27 [1 H, quintet, J 6.8 Hz, NHCH(Me)Ph], 7.17–7.63 (20 H, m, PPh₃ and Ph), 9.98 [1 H, br s, NHCH(Me)Ph]; $\delta_H(CDCl_3)$ 1.54 [3 H, d, J 6.8, NHCH(Me)Ph], 2.52 (3 H, s, Fe=CMe), 4.84 (5 H, d, J_{PH} 1.3 Hz, C₅H₅), 5.00 [1 H, quintet, J 6.8, NHCH(Me)Ph], 6.89–7.53 (20 H, m, PPh₃ and Ph), 9.46 [1 H, br s, NHCH(Me)Ph]; $\delta_C(CDCl_3)$ 22.11 [s, NHCH(Me)Ph], 39.12 (s, Fe=CMe), 60.64 [s, NHCH(Me)Ph], 85.60 (s, C₅H₅), 126.72 (s, C_m of Ph), 128.1 (s, C_p of Ph), 129.1 (s, C_o of Ph), 129.25 (d, J_{PC} 9.8 Hz, C_m of PPh₃), 131.18 (s, C_p of PPh₃), 132.74 (d, J_{PC} 9.7, C_o of PPh₃), 133.17 (d, J_{PC} 44.8, C_{ipso} of PPh₃), 140.38 (s, C_{ipso} of Ph), 217.53 (d, J_{PC} 29.8, Fe–CO), 267.03 (d, J_{PC} 21.9 Hz, Fe=C); $\delta_P(CDCl_3)$ 68.84; $\delta_F(CDCl_3)$ –153.07 and –153.13 (1:3); m/z 558 (M^+ , 71), 530 (M – CO, 92), 383 (8), 268 (100), 239 (9), 226 (14), 203 (13), 183 (32), 162 (20), 146 (13), 121 (20), 105 (49%).**

(*S,S*)-[(η^5 - C_5H_5)Fe(CO)(PPh₃)₂]=C(Me)(NHCH(Me)Ph)]⁺BF₄[–] (S,S**-34). Reaction of (*S*)-6 (0.40 g, 0.72 mmol) with (*S*)- α -methylbenzylamine (0.93 mL, 7.19 mmol) at room temperature for 2 h afforded (*S,S*)-34 as a yellow solid (0.46 g, 95%). Recrystallisation from CH₂Cl₂–EtOH gave yellow plates, m.p. 194–196 °C (decomp.) [lit.,⁸ m.p. 186 °C (decomp.)] (Found: C, 63.42; H, 5.18; N, 2.08; P, 5.03. C₃₄H₃₃BF₄FeNOP requires: C, 63.29; H, 5.15; N, 2.17; P, 4.80%); $[\alpha]_D$ (20 °C, c = 1.005, CH₂Cl₂) 154.6; ν_{max} 1965 (Fe–CO) and 1062 cm^{–1} (BF₄[–]); $\delta_H(CD_3COCD_3)$ 1.41 [3 H, d, J 6.8, NHCH(Me)Ph], 2.72 (3 H, Fe=CMe), 4.97 (5 H, d, J_{PH} 1.4, C₅H₅), 5.32 [1 H, quintet, J 6.8 Hz, NHCH(Me)Ph], 7.29–7.61 (20 H, m, PPh₃ and Ph), 9.90 [1 H, br s, NHCH(Me)Ph]; $\delta_H(CDCl_3)$ 1.40 [3 H, d, J 6.9, NHCH(Me)Ph], 2.48 (3 H, s, Fe=CMe), 4.76 (5 H, d, J_{PH} 1.3, C₅H₅), 4.79 [1 H, quintet, J 6.9 Hz, NHCH(Me)Ph], 7.18–7.53 (20 H, m, PPh₃ and Ph), 9.68 [1 H, br s, NHCH(Me)Ph]; δ_C 21.61 [s, NHCH(Me)Ph], 38.83 (s, Fe=CMe), 60.65 [s, NHCH(Me)Ph], 85.30 (s, C₅H₅), 126.51 (s, C_m of Ph), 128.1 (s, C_p of Ph), 129.1 (s, C_o of Ph), 129.25 (d, J_{PC} 9.8, C_m of PPh₃), 131.18 (s, C_p of PPh₃), 132.74 (d, J_{PC} 9.7 Hz, C_o of PPh₃), 133.08 (d, J_{PC} 44.8, C_{ipso} of PPh₃), 140.60 (s, C_{ipso} of Ph), 217.33 (d, J_{PC} 29.8, Fe–CO), 267.93 (d, J_{PC} 21.9 Hz, Fe=C); $\delta_P(CDCl_3)$ 68.85; $\delta_F(CDCl_3)$ –153.07 and –153.13 (1:3); m/z 558 (M^+ , 60), 530 (M – CO, 13), 383 (8), 268 (100), 239 (9), 226 (12), 183 (19), 163 (19), 105 (20), 89 (23), 77 (35), 59 (85%).**

(*RS,SR*)- and (*RR,SS*)-[(η^5 - C_5H_5)Fe(CO)(PPh₃)₂]=C(Me)(NHCH(Me)Ph)]⁺BF₄[–] (RS,SR**-33 and **RR,SS**-34). A solution of (\pm)- α -methylbenzylamine (4.85 g, 40 mmol) in THF (20 mL) was added to (*RS*)-6 (1.11 g, 2.0 mmol) and the reaction mixture was stirred for 2 h at 0 °C under nitrogen. The yellow precipitate that had formed in the reaction mixture was filtered off, washed with THF, and dried *in vacuo*. The precipitate [0.77 g, 60%, (*RS,SR*):(*RR,SS*) = 92:8] was recrystallised twice from CH₂Cl₂–EtOH to give yellow needles (*RS,SR*) (0.66 g, 51%). The filtrate and washings were combined and concentrated *in vacuo*. The residual oil was distilled at 60–65 °C under vacuum to give (\pm)- α -methylbenzylamine as a colourless oil (2.30 g, 47%) and a yellow residue which was left overnight to crystallise. Filtration followed by washing with benzene gave a yellow solid which consisted (¹H NMR spectroscopy) of (*RS,SR*):(*RR,SS*) in a ratio of 15:85. Column chromatography on alumina (5% MeOH–CH₂Cl₂) followed by two crystallisations**

from CH₂Cl₂-benzene gave the (*RR,SS*) diastereoisomers as yellow needles (0.139 g, 11%).

(*R,S*)- and (*S,S*)-[(η⁵-C₅H₅)Fe(CO)(PPh₃)₂]=C(Me)[NHCH(Me)Ph]⁺BF₄⁻ (*R,S*)-33 and (*S,S*)-34. A solution of (*S*)-(-)-*α*-methylbenzylamine (5.69 g, 47 mmol) in THF (47 mL) was added to (*RS*)-6 (2.64 g, 4.7 mmol) and the reaction mixture was stirred for 2.75 h at 0 °C. The red reaction mixture was concentrated to about 30 mL and was then stirred for a further 1.5 h during which time a precipitate formed. The yellow precipitate was filtered off, washed with THF and dried *in vacuo*. The precipitate (1.14 g, 1.77 mmol, 38%) was shown to consist of a mixture of diastereoisomers of (*R,S*):(*S,S*) in the ratio of 93:7. This solid was recrystallised from EtOH (150 mL) to give pure (*R,S*)-33 as yellow needles (0.95 g, 1.48 mmol, 31%). The filtrate and washings were combined and concentrated *in vacuo*. The residual oil was distilled at 60–65 °C *in vacuo* to give (*S*)-(-)-*α*-methylbenzylamine as a colourless oil {4.32 g, 76%, [α]_D (neat) -39.4°}. The resultant residue was dissolved in a minimum amount of benzene and left to crystallise over 3 d. Filtration followed by washing with benzene gave a mixture of (*R,S*):(*S,S*) in the ratio of 22:78. Two recrystallisations from EtOH gave pure (*S,S*)-34 as yellow plates (0.44 g, 0.68 mmol, 15%).

Coupling reactions with secondary amines and amino acid derivatives

Attempted coupling reactions of (*RS*)-6 with Me₂NH, Et₂NH, pyrrolidine, glycine, glycine methyl ester, and glycine *tert*-butyl ester under the standard reaction conditions afforded mostly unreacted starting material and demethylated (*RS*)-5.

(*RS*)-[(η⁵-C₅H₅)Fe(CO)(PPh₃)₂]=C(Et)(NHMe)]⁺BF₄⁻ (*RS*)-18. Reaction of (*RS*)-17 (1.32 g, 2.37 mmol) with aqueous MeNH₂ (40%, 1.86 mL, 23.7 mmol) for 2 h at 0 °C afforded (*RS*)-18 as a yellow solid (0.89 g, 66%) which was crystallised from CH₂Cl₂-Et₂O as yellow rods (hemisolvate), m.p. 185–189 °C (decomp.) (Found: C, 59.27; H, 5.45; N, 2.31. C₂₈H₂₉BF₄FeNOP·0.5C₄H₁₀O requires: C, 59.44; H, 5.66; N, 2.31%; ν_{max} 1953 (Fe–CO) and 1056 cm⁻¹ (BF₄⁻); δ_H 1.01 (3 H, t, *J* 7.4 Hz, Fe=CCH₂CH₃), 2.18–2.29 (2 H, m, Fe=CCH₂CH₃), 3.03 (3 H, s, NHMe), 4.82 (5 H, s, C₅H₅), 7.31–7.36 (6 H, m, H_o of PPh₃), 7.50–7.55 (9 H, m, H_m and H_p of PPh₃), 9.84 (1 H, br s, NHMe); δ_C 10.15 (s, Fe=CCH₂CH₃), 36.65 (s, NHMe), 41.39 (s, Fe=CCH₂CH₃), 85.68 (s, C₅H₅), 129.19 (d, *J*_{PC} 9.5, C_m of PPh₃), 131.12 (s, C_p of PPh₃), 132.95 (d, *J*_{PC} 8.8, C_o of PPh₃), 133.31 (d, *J*_{PC} 46.0, C_{ipso} of PPh₃), 218.47 (d, *J*_{PC} 30.2, Fe–CO), 268.48 (d, *J*_{PC} 20.1 Hz, Fe=C); δ_F 66.20; δ_F -154.00 and -154.05 (1:3); *m/z* 482 (*M*⁺, 100), 454 (*M* - CO, 27), 383 (13), 318 (9), 295 (46), 263 (23), 220 (9), 192 (98), 183 (27), 85 (29), 70 (34%). The yield of this reaction could not be improved by using gaseous methylamine.

(*RS*)-[(η⁵-C₅H₅)Fe(CO)(PPh₃)₂]=C(Et)(NHEt)]⁺BF₄⁻ (*RS*)-19. Reaction of (*RS*)-17 (1.00 g, 1.80 mmol) with gaseous ethylamine for 2 h at 0 °C afforded (*RS*)-19 as a yellow foam (0.73 g, 70%) which could not be crystallised from a number of solvents. Microanalytical data could not be obtained as this product contained 5% of [(η⁵-C₅H₅)Fe(CO)₂(PPh₃)₂]⁺BF₄⁻ which could not be separated from the product either by chromatography or selective crystallisation; ν_{max} 1954 (Fe–CO) and 1057 cm⁻¹ (BF₄⁻); δ_H 0.85 (3 H, t, *J* 7.1, NCH₂CH₃), 1.08 (3 H, t, *J* 7.6 Hz, Fe=CCH₂CH₃), 2.37–2.56 (2 H, m, Fe=CCH₂CH₃), 3.42–3.51 (2 H, m, NCH₂CH₃), 4.82 (5 H, s, C₅H₅), 7.15–7.65 (15 H, m, PPh₃), 9.61 (1 H, br s, NHEt); δ_C 11.14 (s, Fe=CCH₂CH₃), 13.95 (s, NCH₂CH₃), 41.31 (s, Fe=CCH₂CH₃), 44.62 (s, NCH₂CH₃), 85.34 (s, C₅H₅), 129.11 (d, *J*_{PC} 9.6, C_m of PPh₃), 131.10 (s, C_p of PPh₃), 132.76 (d, *J*_{PC} 8.9, C_o of PPh₃), 133.00 (d, *J*_{PC} 50.5, C_{ipso} of PPh₃), 218.12 (d, *J*_{PC} 28.5, Fe–CO),

267.04 (d, *J*_{PC} 23.2 Hz, Fe=C); δ_P 67.19; δ_F -153.57 and -153.63 (1:3); *m/z* 496 (*M*⁺, 73), 468 (*M* - CO, 21), 383 (12), 318 (10), 295 (10), 279 (32), 263 (13), 206 (100), 183 (17), 129 (40), 84 (27%).

(*RS*)-[(η⁵-C₅H₅)Fe(CO)(PPh₃)₂]=C(Et)(NHCH₂Ph)]⁺BF₄⁻ (*RS*)-20. Reaction of (*RS*)-17 (7.29 g, 0.13 mol) with benzylamine (14.2 mL, 0.13 mol) at 0 °C for 2 h afforded (*RS*)-20 as a yellow solid (6.73 g, 80%). Recrystallisation from CH₂Cl₂-EtOH gave yellow microneedles, m.p. 197–199 °C (decomp.) (Found: C, 63.15; H, 4.91; N, 2.05; P, 4.74. C₃₄H₃₃BF₄FeNOP requires: C, 63.15; H, 5.16; N, 2.17; P, 4.80%; ν_{max} 1957 (Fe–CO) and 1068 cm⁻¹ (BF₄⁻); δ_H 1.09 (3 H, t, *J* 7.6, Fe=CCH₂CH₃), 2.41–2.06 (1 H, m, Fe=CCH₂CH₃), 2.71–2.74 (1 H, m, Fe=CCH₂CH₃), 4.40 (1 H, dd, *J* 14.2, 7.4, NCH₂Ph), 4.84 (1 H, dd, *J* 14.2, 4.5, NCH₂Ph), 4.90 (5 H, d, *J*_{PH} 1.6, C₅H₅), 6.95 (2 H, d, *J* 7.3, H_o of Ph), 7.08 (6 H, dd, *J*_{PH} 10.1, *J*_{HH} 7.5, H_o of PPh₃), 7.09 (3 H, br d, *J* 8.1, H_m of PPh₃), 7.15 (2 H, br t, *J* 7.5, H_m of Ph), 7.21 (1 H, t, *J* 7.5, H_p of Ph), 7.38 (6 H, br t, *J* 7.4, H_m of PPh₃), 7.47 (3 H, br t, *J* 7.9 Hz, H_p of PPh₃), 10.19 (1 H, br s, NHCH₂Ph); δ_C 10.94 (s, Fe=CCH₂CH₃), 41.68 (s, Fe=CCH₂CH₃), 52.08 (s, NCH₂Ph), 85.74 (s, C₅H₅), 128.32 (s, C_p of Ph), 128.69 (s, C_m of Ph), 129.49 (s, C_o of Ph), 129.59 (d, *J*_{PC} 9.7, C_m of PPh₃), 131.59 (s, C_p of PPh₃), 133.26 (d, *J*_{PC} 9.7, C_o of PPh₃), 133.47 (d, *J*_{PC} 45.6, C_{ipso} of PPh₃), 136.75 (s, C_{ipso} of Ph), 218.40 (d, *J*_{PC} 29.8, Fe–CO), 268.45 (d, *J*_{PC} 23.9 Hz, Fe=C); δ_P 67.33; δ_F -153.01 and -153.06 (1:3); *m/z* 558 (*M*⁺, 74), 530 (*M* - CO, 7), 383 (11), 318 (8), 295 (10), 268 (100), 212 (17), 183 (20), 91 (30), 85 (22), 47 (17%).

(*RS*)-[(η⁵-C₅H₅)Fe(CO)(PPh₃)₂]=C(Et)(NHCH₂Ph)]⁺PF₆⁻ (*RS*)-21. A saturated solution of NH₄PF₆ in water (80 mL) was added dropwise to a solution of (*RS*)-20 (1.00 g, 1.55 mmol) in acetone (120 mL) at room temperature under a nitrogen atmosphere. This resulted in the immediate precipitation of a bright yellow solid which was filtered off, washed with water, ethanol and diethyl ether to afford pure (*RS*)-21 (1.04 g, 95%). Recrystallisation from CH₂Cl₂-acetone gave yellow microneedles, m.p. 186–188 °C (Found: C, 58.29; H, 4.72; N, 2.06. C₃₄H₃₃F₆FeNOP₂ requires: C, 58.05; H, 4.73; N, 1.99%; ν_{max} 1958 (Fe–CO) and 842 cm⁻¹ (PF₆⁻); δ_H(CD₃COCD₃) 1.21 (3 H, t, *J* 7.7, Fe=CCH₂CH₃), 2.36–2.42 (1 H, m, Fe=CCH₂CH₃), 3.02–3.07 (1 H, m, Fe=CCH₂CH₃), 4.69 (1 H, dd, *J* 14.7, 4.4, NHCH₂Ph), 4.87 (1 H, dd, *J* 14.7, 6.7, NHCH₂Ph), 5.12 (5 H, d, *J*_{PH} 1.4, C₅H₅), 7.07 (2 H, dd, *J* 7.7, 2.1, H_o of Ph), 7.25 (6 H, dd, *J*_{PH} 10.1, *J*_{HH} 7.7, H_o of PPh₃), 7.31–7.36 (3 H, m, H_m and H_p of Ph), 7.48 (6 H, tq, *J* 7.5, 2.1, H_m of PPh₃), 7.58 (3 H, tq, *J* 7.6, 1.9 Hz, H_p of PPh₃), 10.11 (1 H, br s, NHCH₂Ph); δ_C(CD₃COCD₃) 11.35 (s, Fe=CCH₂CH₃), 43.14 (s, Fe=CCH₂CH₃), 53.82 (s, NCH₂Ph), 86.42 (s, C₅H₅), 128.99 (s, C_p of Ph), 129.25 (s, C_m of Ph), 129.95 (s, C_o of Ph), 129.98 (d, *J*_{PC} 9.2, C_m of PPh₃), 132.00 (s, C_p of PPh₃), 133.70 (d, *J*_{PC} 9.6, C_o of PPh₃), 133.91 (d, *J*_{PC} 43.3, C_{ipso} of PPh₃), 136.25 (s, C_{ipso} of Ph), 218.86 (d, *J*_{PC} 29.1, Fe–CO), 273.03 (d, *J*_{PC} 23.3 Hz, Fe=C); δ_P(CD₃COCD₃) 67.70 (s, PPh₃), -147.11 (septet, *J*_{PF} 708.7 Hz, PF₆⁻); δ_F(CD₃COCD₃) -69.13 (d, *J*_{PF} 707.8 Hz, PF₆⁻); *m/z* 558 (*M*⁺, 87), 530 (*M* - CO, 7), 383 (10), 295 (10), 268 (100), 212 (15), 183 (25), 121 (10), 91 (27), 59 (13%).

(*RS*)-[(η⁵-C₅H₅)Fe(CO)(PPh₃)₂]=C(CH₂CH₂CH₂CH₃)(NHMe)]⁺BF₄⁻ (*RS*)-24. Reaction of (*RS*)-23 (1.18 g, 1.97 mmol) with gaseous methylamine at 0 °C for 2 h afforded (*RS*)-24 as a yellow foam (0.88 g, 80%). Microanalytical data could not be obtained for this product which was isolated as a mixture with [(η⁵-C₅H₅)Fe(CO)₂(PPh₃)₂]⁺BF₄⁻ (10%). This impurity could not be removed by column chromatography or selective recrystallisation; ν_{max} 1964 (Fe–CO) and 1054 cm⁻¹ (BF₄⁻); δ_H 0.88 (3 H, t, *J* 7.1 Hz, Fe=CCH₂CH₂CH₂CH₃), 1.10–2.50 (6 H, m, Fe=CCH₂CH₂CH₂CH₃), 3.00 (3 H, s, NMe), 4.79 (5 H, s, C₅H₅), 7.19–7.82 (15 H, m, PPh₃), 10.19 (1 H, br s, NHMe);

δ_C 13.46 (s, Fe=CCH₂CH₂CH₂CH₃), 22.66 (s, Fe=CCH₂CH₂CH₂CH₃), 27.75 (s, Fe=CCH₂CH₂CH₂CH₃), 36.12 (s, NMe), 48.05 (s, Fe=CCH₂CH₂CH₂CH₃), 85.33 (s, C₅H₅), 129.04 (d, J_{PC} 9.8, C_m of PPh₃), 131.06 (s, C_p of PPh₃), 133.82 (d, J_{PC} 9.8, C_o of PPh₃), 133.11 (d, J_{PC} 46.3, C_{ipso} of PPh₃), 218.28 (d, J_{PC} 28.5, Fe-CO), 267.94 (d, J_{PC} 23.5 Hz, Fe=C); δ_F 66.25; δ_F -153.85 and -153.90 (1:3); m/z 510 (M^+ , 44), 482 (M - CO, 18), 220 (100), 183 (20), 129 (72), 98 (40%).

Deuteration studies of (RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₂]=C(Me)-(NHCH₂Ph)]⁺PF₆⁻ (RS)-15

Method (a). Treatment of a nitrogen degassed solution of **15** (50 mg) in CD₃OD with NaOMe (20 mol %) at room temperature for 3 d afforded by ¹H NMR analysis (RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₂]=C(CD₃)(NDCH₂Ph)]⁺PF₆⁻.

Method (b). Treatment of a solution of (RS)-**15** (100 mg) in THF (10 mL) with BuLi (1 mol equivalent) at -78 °C for 30 min followed by quenching with CD₃OD and concentration *in vacuo* afforded by ¹H NMR analysis (RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₂]=C(Me)(NDCH₂Ph)]⁺PF₆⁻ (RS)-**28**.

Method (c). Treatment of a solution of (RS)-**15** (100 mg) in THF (10 mL) with BuLi (3 or more mol equivalents) at -78 °C for 3 h followed by quenching with either CD₃OD or D₂O and filtration through dry Celite afforded (RS)-[(η^5 -C₅H₅)Fe(CO)(PPh₃)₂]=C(CD₂H)(NDCH₂Ph)]⁺PF₆⁻ (RS)-**30** which was recrystallised from CH₂Cl₂-hexane as yellow microneedles, m.p. 197–199 °C (decomp.) (Found: C, 57.40; H, 4.80; N, 2.03. C₃₃H₂₈BD₃F₄FeNOP requires: C, 57.24; H, 4.95; N, 2.02%); ν_{max} 1958 (Fe-CO) and 841 cm⁻¹ (PF₆⁻); δ_H (CD₂Cl₂) 2.69 (1 H, br s, Fe=CD₂H), 4.53 (1 H, dd, J 14.6, 4.8, NHCH₂Ph), 4.58 (1 H, dd, J 14.4, 5.9, NHCH₂Ph), 4.77 (5 H, s, C₅H₅), 7.00 (2 H, br d, J 7.3, H_o of Ph), 7.10 (6 H, dd, J_{PH} 10.8, J_{HH} 7.7, H_o of PPh₃), 7.30–7.38 (3 H, m, H_m and H_p of Ph), 7.41 (6 H, td, J 7.6, 1.9, H_m of PPh₃), 7.51 (3 H, tq, J 7.5, 1.7 Hz, H_p of PPh₃); m/z 547 (M^+ , 68), 519 (M - CO, 10), 383 (18), 338 (17), 318 (11), 295 (31), 279 (32), 263 (37), 256 (100), 212 (35), 183 (25), 161 (30), 148 (13), 133 (15), 121 (17), 103 (24), 91 (41%).

Crystal structure determination of (R,S)-**33**

C₃₄H₃₃BF₄FeNOP, $M = 589.42$, orthorhombic, $P2_21_21$ (no. 18), $a = 10.603(1)$, $b = 11.894(1)$, $c = 24.779(3)$ Å (from least squares fitting of setting angles for 25 reflections $20.1 \leq \theta \leq 32.9^\circ$), $U = 3125$ Å³, $Z = 4$, $D_x = 1.253$ g cm⁻³, Cu-K α radiation, orange needles $0.2 \times 0.2 \times 0.1$ mm, $\mu = 11.99$ cm⁻¹. 7271 Reflections measured, 5759 unique ($R_{merge} = 0.03$, Friedel pairs were not merged) of which 2978 were observed ($I \geq 3\sigma(I)$). At con-

vergence $R = 0.038$, $R' = 0.031$ for 399 parameters. CCDC reference number 186/949.

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References

- 1 S. G. Davies, *Pure Appl. Chem.*, 1988, **60**, 40.
- 2 S. G. Davies, *Aldrichimica Acta*, 1990, **23**, 31.
- 3 S. G. Davies, I. M. Dordor-Hedgecock, K. H. Sutton and M. Whittaker, *J. Am. Chem. Soc.*, 1987, **109**, 5711.
- 4 G. J. Baird, S. G. Davies and T. R. Maberly, *Organometallics*, 1984, **3**, 1764.
- 5 S. G. Davies and T. R. Maberly, *J. Organomet. Chem.*, 1985, **296**, C37.
- 6 G. J. Baird, S. G. Davies, R. H. Jones, K. Prout and P. Warner, *J. Chem. Soc., Chem. Commun.*, 1984, 744.
- 7 L. S. Hegedus, M. A. Schwindt, S. D. Lombaert and R. Imwinkelried, *J. Am. Chem. Soc.*, 1990, **112**, 2264.
- 8 A. Davison and D. L. Reger, *J. Am. Chem. Soc.*, 1972, **94**, 9237.
- 9 (a) S. Abbott, S. G. Davies and P. Warner, *J. Organomet. Chem.*, 1983, **246**, C65; (b) B. E. Boland-Lussier and R. P. Hughes, *Organometallics*, 1982, **1**, 635.
- 10 (a) M. L. H. Green, L. C. Mitchard and M. G. Swanick, *J. Chem. Soc. A*, 1971, 794; (b) T. Bodnar and A. R. Cutler, *J. Organomet. Chem.*, 1981, **213**, C31; (c) M. Brockhard, J. R. Tucker and G. R. Husk, *J. Am. Chem. Soc.*, 1983, **105**, 258.
- 11 B. J. Wakefield, *Organolithium Methods*, Academic Press, London, 1988.
- 12 S. C. Case-Green, J. F. Costello, S. G. Davies, N. Heaton, C. J. R. Hedgecock, V. M. Humphreys, M. R. Metzler and J. C. Prime, *J. Chem. Soc., Perkin Trans. 1*, 1994, 933.
- 13 M. Brookhart, Y. Liu, W. Goldman, D. A. Timmers and G. D. Williams, *J. Am. Chem. Soc.*, 1991, **113**, 927.
- 14 B. K. Blackburn, S. G. Davies and M. Whittaker, *Chemical Bonds – Better Ways to Make Them and Break Them*, ed. I. Bernal, Oxford, 1989, vol. 3, pp. 141–223.
- 15 S. G. Davies, M. R. Metzler, K. Yanada and R. Yanada, *J. Chem. Soc., Chem. Commun.*, 1993, 658.
- 16 D. F. Shriver, *The Manipulation of Air-Sensitive Compounds*, McGraw-Hill, New York, 1969.
- 17 J. P. Bibler and A. Wojcicki, *Inorg. Chem.*, 1966, **5**, 889.
- 18 M. Green and D. J. Westlake, *J. Chem. Soc. A*, 1971, 367.
- 19 S. Abbott, G. J. Baird, S. G. Davies, I. M. Dordor-Hedgecock, T. R. Maberly, J. C. Walker and P. Warner, *J. Organomet. Chem.*, 1985, **289**, C13.
- 20 S. G. Davies, M. R. Metzler, W. C. Watkins, R. G. Compton, J. Booth and J. C. Eklund, *J. Chem. Soc., Perkin Trans. 2*, 1993, 1603.
- 21 (a) A. C. Cope, C. R. Ganellin, H. W. Johnson, jun., T. V. Van Auken and H. J. S. Winkler, *J. Am. Chem. Soc.*, 1963, **85**, 3276; (b) D. Parker and R. J. Taylor, *Tetrahedron*, 1987, **43**, 5451.

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